

Why is a Noble Metal Catalytically Active? The Role of the O-Ag Interaction in the Function of Silver as an Oxidation Catalyst

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(Received 11 August 2002; published 25 June 2003)

Extensive density-functional theory calculations, and taking into account temperature and pressure, afford a comprehensive picture of the behavior and interaction of oxygen and Ag(111), and provides valuable insight into the function of silver as an oxidation catalyst. The obtained phase diagram reveals the most stable species present in a given environment and thus identifies (and excludes) possibly active oxygen species. In particular, for the conditions of ethylene epoxidation, a thin oxidelike structure is most stable, suggesting that such atomic O species are actuating the catalysis, in contrast to hitherto proposed molecularlike species.

DOI: 10.1103/PhysRevLett.90.256102

PACS numbers: 81.65.Mq, 68.43.Bc

The importance of reactant bond strength in heterogeneous catalysis is demonstrated by the so-called “volcano curves” [1–3] which relate the reaction rate to the adsorption energy of the reaction intermediates. Typically a maximum occurs in the rate at moderate values. This behavior underlies the expectation that a good catalyst should readily dissociate adparticles but not bind the fragments too strongly—a concept put forward by Sabatier in the early 1900s, which recent theoretical and experimental studies have readdressed [4,5]. Despite this expectation, silver, as a noble metal, which binds adparticles only relatively weakly on the surface, is a very important heterogeneous catalyst for various oxidation reactions. For example, partial oxidation of methanol to formaldehyde, as carried out at atmospheric pressure and temperatures (T) of 800–900 K [6,7] as well as the selective oxidation of ethylene to epoxide, also conducted at atmospheric pressure and at $T = 500$ –600 K [8].

Although numerous investigations aimed at elucidating the active oxygen species involved in the above-mentioned reactions have been carried out, it can be said that presently there is considerable confusion, and no unambiguous identifications have been made. For example, unclear issues concern the debate of molecular (ozonelike) [9] versus atomic [10] oxygen as the active species for ethylene epoxidation, the proposed importance of “bulk-dissolved” oxygen and a “strongly bound, surface-embedded” species (which desorbs at 900 K) for the partial oxidation of methanol [6], as well as the possible role of chemisorbed or surface oxidelike species. Part of the reason for this lack of understanding is due to the difficulties associated with the pressure and materials gaps which exist for the oxygen-silver system. That is, the behavior is different under ultrahigh vacuum (UHV) conditions, where the system can be analyzed quantitatively, to that under the high temperature and pressure conditions of catalysis, where it is much more difficult to obtain the same level of microscopic information.

Our theoretical strategy, aimed at gaining an understanding of the function of silver as an oxidation catalyst, in spite of its weak binding capability, and identifying (and excluding) possibly active species, is to carry out systematic density-functional theory (DFT) calculations for all conceivably relevant oxygen species. We take the effect of temperature T and pressure p into account through the oxygen chemical potential, and thus determine the (T, p) phase diagram which describes the surface phases from UHV right up to the conditions of real catalysis. We study oxygen at the (111) surface of silver as experiments indicate that this orientation is an important crystal face for real silver catalysts in that at high temperatures, facets with this face result [6,11]. Our general findings, however, are expected to be relevant for silver *per se*, and possibly to have implications for gold, also a noble metal oxidation catalyst [12] which, for the (111) surface, exhibits a restructuring and chemisorption of oxygen atoms at elevated temperatures (500–800 K) and atmospheric pressures [13,14], similar to Ag(111). They could also have consequences for copper, which catalyzes the oxidation of methanol to formaldehyde [15].

Our calculations employ the pseudopotential [16,17] plane-wave method [18] with the generalized gradient approximation (GGA) [19] for the exchange-correlation functional. Five Ag layers are used in the supercell to model the O/Ag surfaces, which are created on one side of the slab, with a vacuum region of 15 Å. The same \mathbf{k} -point sampling is used for all structures. It corresponds to 21 points in the irreducible part of the Brillouin zone of the (1×1) Ag(111) surface cell. An energy cutoff of 50 Ry is used and full relaxation of the top two or three silver layers, and the O atoms, is taken into account. We include the spin-polarization energy in calculation of the total energy of the free atoms and molecules. Further details are reported in Ref. [20].

The oxygen species investigated, for a wide range of coverages, include on-surface sites, surface- and bulk-substitutional sites, interstitial sites under the surface Ag

layer and deeper in the bulk, as well as oxidelike structures, and a molecular ozonelike species adsorbed at a surface Ag vacancy.

The main important adsorption energies are summarized in Fig. 1, where for each “type” of oxygen atom, only the lowest energy structure is shown. For on-surface adsorption, the fcc-hollow site is preferred for all coverages investigated. For subsurface adsorption, the octahedral site under the first Ag layer is the most favorable. This is also the case for adsorption under the second Ag layer which may be taken to represent “bulk dissolved” oxygen. For all subsurface sites, adsorption under the second layer is *less* favorable than under the first layer. In view of the unfavorable energy of oxygen in the bulk, high concentrations of bulk-dissolved oxygen in the perfect fcc lattice of silver are improbable.

Surface- or bulk-substitutional adsorption is seen from Fig. 1 to be unfavorable, as is the ozonelike species. To investigate the formation of oxidelike structures, we performed systematic investigations for geometries involving coverages of 0.50, 0.75, 1.0, and 1.25 ML using (2×2)

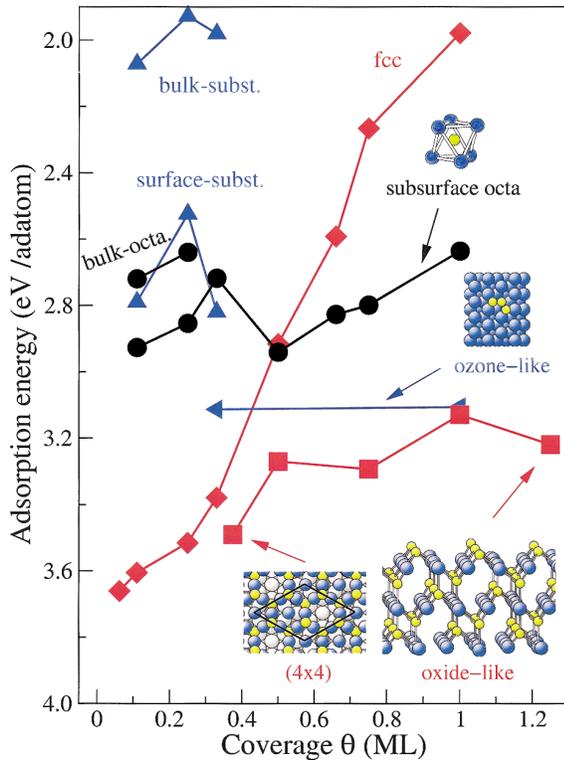


FIG. 1 (color). Selected results of average adsorption energies (with respect to atomic O) versus coverage for various structures: pure on-surface oxygen in fcc-hollow sites (red diamonds), surface- and bulk-substitutional adsorption (blue triangles), pure subsurface oxygen in octahedral sites under the first (“subsurface-octa.”) and second (“bulk-octa.”) Ag layers (black circles), an O_3 -like molecule adsorbed at a surface Ag vacancy (left-pointing black triangles), and oxide-like structures (red squares). Oxygen atoms are depicted as yellow spheres and Ag atoms as the larger blue (and grey) spheres.

cells. In particular, we tested combinations of on-surface fcc and hcp sites, and the three possible subsurface sites under the first Ag layer. The atomic geometry of the favored structure consists of 0.25 ML of oxygen on the surface and 0.5 ML in between the Ag layers. For increasing O concentrations, this structure “grows” by filling deeper lying sites between the Ag layers. The structure for a total coverage of 1.25 ML oxygen is sketched in the inset of Fig. 1. In addition, we carried out calculations for the recently proposed (4×4) thin “surface-oxide” structure [21] which involves an O-Ag-O bonding unit (and electronic structure) similar to the upper trilayer of the oxide structures described above (see insets of Fig. 1). The average adsorption energies of all these oxidelike structures are displayed in Fig. 1, shown as red squares. Their atomic geometries are, in fact, very similar to the (111) surface of Ag_2O .

We turn now to investigate the effect of pressure and temperature on the stability of the various structures. To do this we calculate the surface free energy,

$$\gamma(T, p) = (G - N_{Ag}\mu_{Ag} - N_O\mu_O)/A, \quad (1)$$

where $G = G_{O/Ag(111)}^{slab} - G_{Ag(111)}^{slab}$ and the first and second terms on the right-hand side are the free energies of the O/Ag surface under consideration and that of the reference system, i.e., the clean Ag(111) slab, respectively. A is surface area in \AA^2 and N_O is the number of oxygen atoms. N_{Ag} is the number of silver atoms (with respect to the reference system) and μ_{Ag} is the Ag chemical potential, which is the free energy of an Ag atom in bulk fcc silver. [Consideration of N_{Ag} and μ_{Ag} is necessary only for the substitutional structures and the (4×4) phase.] The T and p dependence is given by μ_O , the oxygen chemical potential [22]

$$\mu_O(T, p) = 1/2 \left[E_{O_2}^{total} + \tilde{\mu}_{O_2}(T, p^0) + k_B T \ln \left(\frac{p_{O_2}}{p^0} \right) \right], \quad (2)$$

where p^0 corresponds to atmospheric pressure and $\tilde{\mu}_{O_2}(T, p^0)$ includes the contribution from rotations and vibrations of the molecule, as well as the ideal-gas entropy at 1 atm. Here we use experimental values from thermodynamic tables.

Through consideration of the difference in the vibrational and entropic contributions of $G_{O/Ag(111)}^{slab}$ and $G_{Ag(111)}^{slab}$, we find that it can be neglected [23]. Thus, for evaluating *the difference* of the two G^{slab} values, we replace G^{slab} by the total energy. We choose the zero of the O chemical potential to be half the total energy of O_2 calculated as $1/2 E_{O_2}^{total} \approx E_{Ag_2O}^{total} - 2E_{Ag}^{total} - H_{Ag_2O-bulk}^f$, that is, as the energy difference between bulk silver oxide ($E_{Ag_2O}^{total}$) and bulk silver ($2E_{Ag}^{total}$), and the experimental value of the heat of formation ($H_{Ag_2O-bulk}^f$) which is -0.323 eV [24].

The obtained results are displayed in Fig. 2. The range of $\Delta\mu_{\text{O}}$ (defined as $\mu_{\text{O}} - \frac{1}{2}E_{\text{O}_2}^{\text{total}}$) between the vertical dashed lines corresponds to the heat of formation of Ag_2O , i.e., to the range in which bulk silver oxide is stable. It can be seen that for more O-rich conditions (right) the thicker oxidelike structures are favored over chemisorption on the surface. It can be obtained from Fig. 2 that bulk silver oxide is stable to only around 350 K at atmospheric pressure, which is less than the experimental value of 460 K. At 460 K, the bulk oxide is unstable when the pressure is lower than 10^3 atm, i.e., compared to 1 atm as observed experimentally. This difference may be due to *systematic* errors of the DFT approach [25] and/or to neglect of the entropy contributions. Therefore in Fig. 2 (and Fig. 3), we generally expect that the temperature is underestimated, but the error is less than 110 K, while correspondingly, the pressure is overestimated. For the temperature values quoted below, we give an estimate of the upper limit in brackets, that is, 110 K higher than that

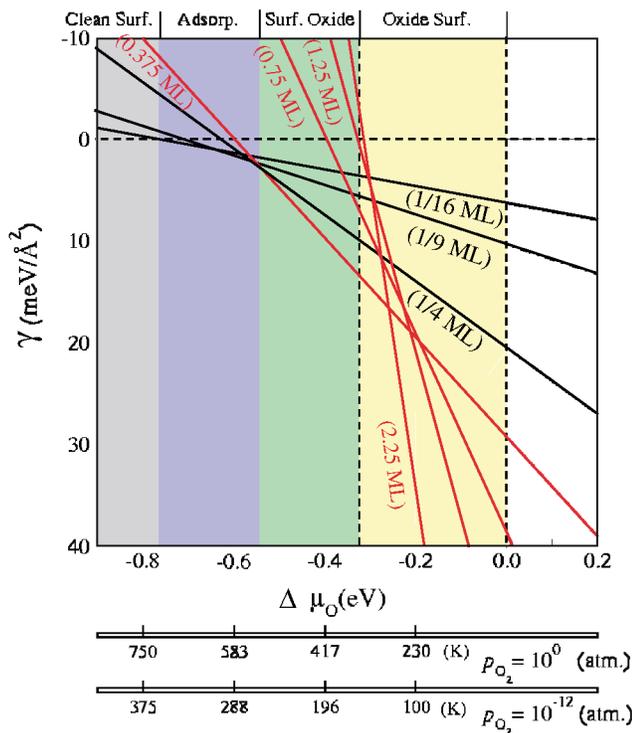


FIG. 2 (color). Surface free energies for various low energy structures as a function of the O chemical potential where $\Delta\mu_{\text{O}} = \mu_{\text{O}} - \frac{1}{2}E_{\text{O}_2}^{\text{total}}$. For pressures of 10^{-12} atm (UHV) and 1 atm (condition of catalysis), the corresponding temperatures are given. The temperature labels correspond to the tick marks and labels for the chemical potential. The labels “0.75 ML,” “1.25 ML,” and “2.25 ML” indicate the O concentrations in the corresponding oxidelike structures. Accordingly, the (4×4) structure has the label 0.375 ML, and the adsorption structures 1/4, 1/9, and 1/16 ML. At the top of the figure, the “material type” which is stable in the corresponding range of chemical potential is listed and indicated by the colored regions.

obtained from Fig. 2. In spite of these uncertainties, we believe that the relative stability of the various systems and our general understanding and conclusions will not be affected.

Because of the above-mentioned low thermal stability of silver oxide and the thicker oxidelike structures, our results can safely rule them out as playing an important role in the catalytic reactions. For values of the O chemical potential farther to the left in Fig. 2, and atmospheric pressure, the (4×4) phase is the most stable for the temperature range 350–530 K (including corrections, 460–640 K). For higher temperatures, up to about 720 K (including corrections, 830 K), on-surface adsorbed oxygen is the only stable species. For temperatures beyond this, there are no stable species except for O atoms adsorbed at undercoordinated surface Ag atoms, such as next to a vacancy or step edge, which are stable to 865 K (including corrections, 975 K). From the above analysis, our results point to oxygen atoms of the thin (4×4) surface-oxide structure as the main species actuating ethylene epoxidation as this occurs experimentally at $p_{\text{O}_2} = 1$ atm and $T = 500$ –600 K, while either low coverages of on-surface oxygen or oxygen adsorbed at defects may play a role in the higher temperature catalytic reactions.

In Fig. 3, we consider just the lowest energy structures as a function of T and p , yielding the phase diagram. We note that the phase diagram is in-line with available well-established experimental results. Figure 3 furthermore shows that regardless of the pressure, the ordering of the phases is the same. Thus, the same phase should be observable at different pressures if the temperature is adjusted accordingly, which indicates a way to bridge

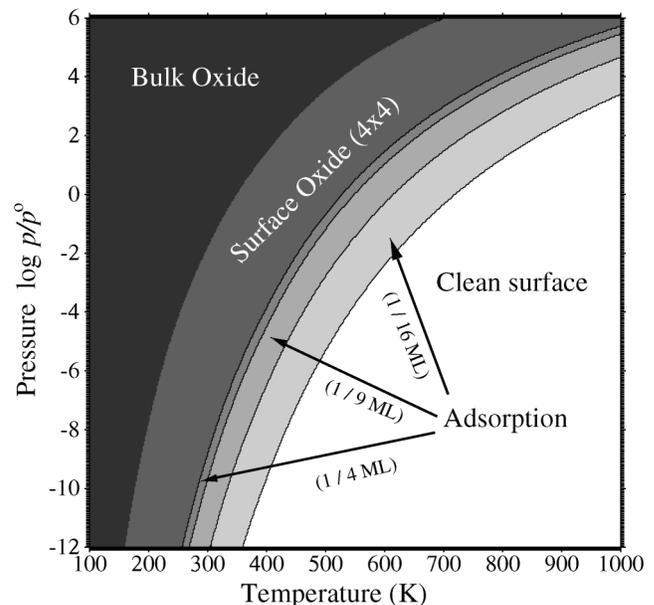


FIG. 3. Calculated (p, T) phase diagram for the oxygen-Ag(111) system showing the stable structures.

the pressure gap—providing that formation of the structure is not prevented by kinetics.

It is interesting to consider the “removal” energy of a single (uppermost) O atom of the (4×4) structure, which is the energy required to move the O atom into the vacuum. This energy is the strongest that we found for the O/Ag system, namely, ~ 3.9 eV, which is due to the presence of the lower lying O atoms which stabilizes the upper O atoms. Compared to other O/metal systems, however, this does not constitute a strong bond. In fact, it represents an *intermediate* value, very similar to the identified “optimum” values for highest catalytic activity of certain reactions [4]. Interestingly, recent studies of the CO oxidation reaction over Ru have surprisingly found that oxides play an active role, in contrast to the hitherto believed pure metal [26,27]. Ruthenium is a poor oxidation catalyst since it binds O too strongly, but at the high T and p conditions where Ru oxides form, a new weaker O species appears and the system is then catalytically very active. Thus it appears that the interaction of oxygen with transition metal catalysts can serve to “tune” the O-metal bond strength for optimum reactivity through formation of various oxidelike species.

In conclusion, through DFT calculations, and including the effect of the environment, we obtained the pressure–temperature phase diagram for O/Ag(111). The results show that a proposed molecular ozonelike species adsorbed at a surface vacancy is energetically *unfavorable*, as is bulk-dissolved oxygen, which puts a question mark over the latter’s hitherto thought important role in oxidation reactions. Our results also show that silver oxide and oxidelike films can also be ruled out as playing an important role due to their low thermal stability. On the other hand, the calculations reveal that silver can support a *thin* surface oxidelike structure under the temperature and pressure conditions of ethylene epoxidation and we propose that such atomic O species are actuating the reaction. For even higher temperatures, only very low concentrations of chemisorbed O are stable, or O atoms adsorbed at undercoordinated surface Ag atoms. Therefore these species could possibly play a role in the catalytic oxidation reactions that take place at high temperature. We note that an understanding of a *full catalytic cycle* requires a kinetic modeling that includes all reactant species and intermediates, which, for complex systems such as those discussed here, is not yet possible. The present investigation represents a crucial first step towards this aim, which, using the approach of atomistic *ab initio* thermodynamics, identifies (and excludes) oxygen species that are present under the conditions of catalysis, and helps shed new light upon the function of silver as an oxidation catalyst.

More generally, our results highlight the formation and importance of various oxides under catalytic conditions, be it in weaker bonded O/noble-metal systems, such as O/Ag, or in stronger bonded O/transition-metal systems,

as, e.g., O/Ru. For transition metals that form bulk oxides with intermediate heats of formation, e.g., PdO, under conditions typical of catalysis, we may expect that both bulk oxides and thinner surface oxides could occur, depending sensitively on the pressure and temperature.

We thank D. King and R. Schlögl for helpful discussions.

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